OXFORD
INTERNATIONAL
AQA EXAMINATIONS

INTERNATIONAL AS Chemistry

(9620)

CH01: Inorganic 1 and Physical 1
Mark Scheme

January 2018

Version: 1.0 Final

Mark schemes are prepared by the Lead Assessment Writer and considered, together with the relevant questions, by a panel of subject teachers. This mark scheme includes any amendments made at the standardisation events which all associates participate in and is the scheme which was used by them in this examination. The standardisation process ensures that the mark scheme covers the students' responses to questions and that every associate understands and applies it in the same correct way. As preparation for standardisation each associate analyses a number of students' scripts. Alternative answers not already covered by the mark scheme are discussed and legislated for. If, after the standardisation process, associates encounter unusual answers which have not been raised they are required to refer these to the Lead Assessment Writer.

It must be stressed that a mark scheme is a working document, in many cases further developed and expanded on the basis of students' reactions to a particular paper. Assumptions about future mark schemes on the basis of one year's document should be avoided; whilst the guiding principles of assessment remain constant, details will change, depending on the content of a particular examination paper.

A-level Chemistry
Mark Scheme Instructions for Examiners

1. General

The mark scheme for each question shows:

- the marks available for each part of the question
- the total marks available for the question
- the typical answer or answers which are expected
- extra information to help the examiner make his or her judgement and help to delineate what is acceptable or not worthy of credit or, in discursive answers, to give an overview of the area in which a mark or marks may be awarded.

The extra information in the 'Comments' column is aligned to the appropriate answer in the left-hand part of the mark scheme and should only be applied to that item in the mark scheme.

You should mark according to the contents of the mark scheme. If you are in any doubt about applying the mark scheme to a particular response, consult your Team Leader.

At the beginning of a part of a question a reminder may be given, for example: where consequential marking needs to be considered in a calculation; or the answer may be on the diagram or at a different place on the script.

In general the right-hand side of the mark scheme is there to provide those extra details which might confuse the main part of the mark scheme yet may be helpful in ensuring that marking is straightforward and consistent.

The use of M1, M2, M3 etc in the right-hand column refers to the marking points in the order in which they appear in the mark scheme. So, M1 refers to the first marking point, M2 the second marking point etc.

2. Emboldening

2.1 In a list of acceptable answers where more than one mark is available 'any **two** from' is used, with the number of marks emboldened. Each of the following bullet points is a potential mark.

2.2 A bold **and** is used to indicate that both parts of the answer are required to award the mark.

2.3 Alternative answers acceptable for a mark are indicated by the use of **OR**. Different terms in the mark scheme are shown by a / ; eg allow smooth / free movement.

3. Marking points

3.1 Marking of lists

This applies to questions requiring a set number of responses, but for which students have provided extra responses. The general 'List' principle to be followed in such a situation is that 'right + wrong = wrong'.

Each error / contradiction negates each correct response. So, if the number of error / contradictions equals or exceeds the number of marks available for the question, no marks can be awarded.

However, responses considered to be neutral (often prefaced by 'Ignore' in the mark scheme) are not penalised.

For example, in a question requiring 2 answers for 2 marks:

Correct answers	Incorrect answers (i.e. incorrect rather than neutral)	Mark (2)	Comment
1	0	1	
1	1	1	They have not exceeded the maximum number of responses so there is no penalty.
1	2	0	They have exceeded the maximum number of responses so the extra incorrect response cancels the correct one.
2	0	2	
2	1	1	
2	2	0	
3	0	2	The maximum mark is 2
3	1	1	The incorrect response cancels out one of the two correct responses that gained credit.
3	2	0	Two incorrect responses cancel out the two marks gained.
3	3	0	

3.2 Marking procedure for calculations

Full marks should be awarded for a correct numerical answer, without any working shown, unless the question states 'Show your working' or 'justify your answer'. In this case, the mark scheme will clearly indicate what is required to gain full credit.

If an answer to a calculation is incorrect and working is shown, process mark(s) can usually be gained by correct substitution / working and this is shown in the 'Comments' column or by each stage of a longer calculation.

3.3 Errors carried forward, consequential marking and arithmetic errors

Allowances for errors carried forward are most likely to be restricted to calculation questions and should be shown by the abbreviation ECF or consequential in the marking scheme.

An arithmetic error should be penalised for one mark only unless otherwise amplified in the marking scheme. Arithmetic errors may arise from a slip in a calculation or from an incorrect transfer of a numerical value from data given in a question.

3.4 Equations

In questions requiring students to write equations, state symbols are generally ignored unless otherwise stated in the 'Comments' column.

Examiners should also credit correct equations using multiples and fractions unless otherwise stated in the 'Comments' column.

3.5 Oxidation states

In general, the sign for an oxidation state will be assumed to be positive unless specifically shown to be negative.

3.6 Interpretation of 'it'

Answers using the word 'it' should be given credit only if it is clear that the 'it' refers to the correct subject.

3.7 Phonetic spelling

The phonetic spelling of correct scientific terminology should be credited **unless** there is a possible confusion with another technical term or if the question requires correct IUPAC nomenclature.

3.8 Brackets

(…..) are used to indicate information which is not essential for the mark to be awarded but is included to help the examiner identify the sense of the answer required.

3.9 Ignore / Insufficient / Do not allow

Ignore or insufficient is used when the information given is irrelevant to the question or not enough to gain the marking point. Any further correct amplification could gain the marking point.

Do **not** allow means that this is a wrong answer which, even if the correct answer is given, will still mean that the mark is not awarded.

3.10 Marking crossed out work

Crossed out work that **has not been** replaced should be marked as if it were not crossed out, if possible. Where crossed out work **has been** replaced, the replacement work and not the crossed out work should be marked.

3.11 Reagents

The command word "Identify", allows the student to choose to use **either** the name or the formula of a reagent in their answer. In some circumstances, the list principle may apply when both the name and the formula are used. Specific details will be given in mark schemes.

The guiding principle is that a reagent is a chemical which can be taken out of a bottle or container. Failure to identify complete reagents **will be penalised**, but follow-on marks (e.g. for a subsequent equation or observation) can be scored from an incorrect attempt (possibly an incomplete reagent) at the correct reagent. Specific details will be given in mark schemes.

For example, **no credit** would be given for

- the cyanide ion or CN^- when the reagent should be potassium cyanide or KCN;
- the hydroxide ion or OH^- when the reagent should be sodium hydroxide or NaOH;
- the $Ag(NH_3)_2^+$ ion when the reagent should be Tollens' reagent (or ammoniacal silver nitrate). In this example, no credit is given for the ion, but credit could be given for a correct observation following on from the use of the ion. Specific details will be given in mark schemes.

In the event that a student provides, for example, **both** KCN and cyanide ion, it would be usual to ignore the reference to the cyanide ion (because this is not contradictory) and credit the KCN. Specific details will be given in mark schemes.

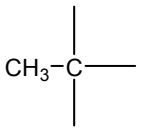
3.12 Organic structures

Where students are asked to draw organic structures, unless a specific type is required in the question and stated in the mark scheme, these may be given as displayed, structural or skeletal formulas or a combination of all three as long as the result is unambiguous.

In general

- Displayed formulae must show all of the bonds and all of the atoms in the molecule, but need not show correct bond angles.
- Skeletal formulae must show carbon atoms by an angle or suitable intersection in the skeleton chain. Functional groups must be shown and it is essential that all atoms other than C atoms are shown in these (except H atoms in the functional groups of aldehydes, secondary amines and N-substituted amides which do not need to be shown).
- Structures must not be ambiguous, e.g. 1-bromopropane should be shown as $CH_3CH_2CH_2Br$ and not as the molecular formula C_3H_7Br which could also represent the isomeric 2-bromopropane.
- Bonds should be drawn correctly between the relevant atoms. This principle applies in all cases where the attached functional group contains a carbon atom, e.g nitrile, carboxylic acid, aldehyde and acid chloride. The carbon-carbon bond should be clearly shown. Wrongly bonded atoms will be penalised **on every occasion**. (see the examples below)
- The same principle should also be applied to the structure of alcohols. For example, if students show the alcohol functional group as C ─ HO, they should be penalised **on every occasion**.
- Latitude should be given to the representation of C ─ C bonds in alkyl groups, given that CH_3─ is considered to be interchangeable with H_3C─ even though the latter would be preferred.
- Similar latitude should be given to the representation of amines where NH_2─ C will be allowed, although H_2N─ C would be preferred.
- Poor presentation of vertical C ─ CH_3 bonds or vertical C ─ NH_2 bonds should **not** be penalised. For other functional groups, such as ─ OH and ─ CN, the limit of tolerance is the half-way position between the vertical bond and the relevant atoms in the attached group.

By way of illustration, the following would apply.



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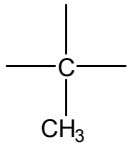

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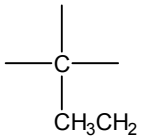

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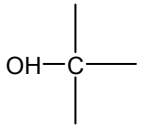

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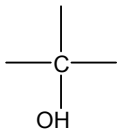

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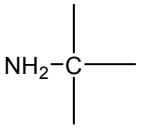

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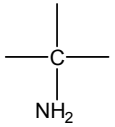

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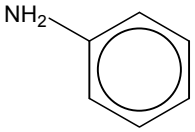

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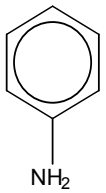

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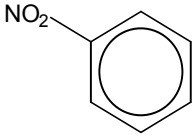

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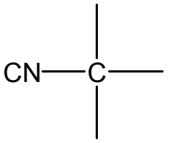

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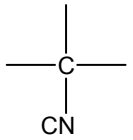

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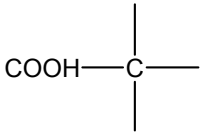

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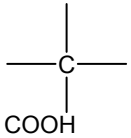

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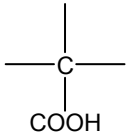

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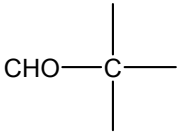

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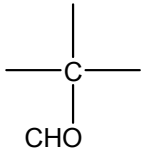

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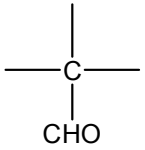

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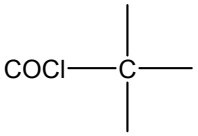

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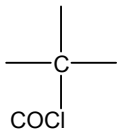

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- Representation of CH_2 by $C–H_2$ will be penalised
- Some examples are given here of **structures** for specific compounds that should **not** gain credit (but, exceptions may be made in the context of balancing equations)

CH_3COH	for	ethanal
CH_3CH_2HO	for	ethanol
$OHCH_2CH_3$	for	ethanol
C_2H_6O	for	ethanol
CH_2CH_2	for	ethene
$CH_2.CH_2$	for	ethene
$CH_2:CH_2$	for	ethane

- Each of the following **should gain credit** as alternatives to correct representations of the structures.

$CH_2 = CH_2$	for	ethene, $H_2C=CH_2$
$CH_3CHOHCH_3$	for	propan-2-ol, $CH_3CH(OH)CH_3$

- In most cases, the use of “sticks” to represent C ─ H bonds in a structure should **not** be penalised. The exceptions to this when “sticks” will be penalised include
  - structures in mechanisms where the C ─ H bond is essential (e.g. elimination reactions in halogenoalkanes and alcohols)
  - when a displayed formula is required
  - when a skeletal structure is required or has been drawn by the candidate

3.13 Organic names

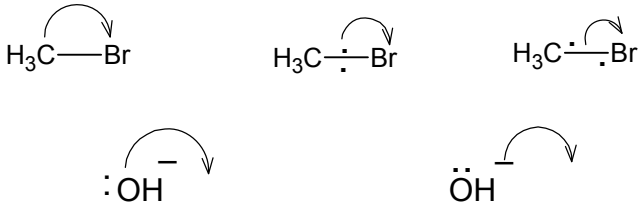
As a general principle, non-IUPAC names or incorrect spelling or incomplete names should **not** gain credit. Some illustrations are given here.

but-2-ol	should be **butan-2-ol**
2-hydroxybutane	should be **butan-2-ol**
butane-2-ol	should be **butan-2-ol**
2-butanol	should be **butan-2-ol**
ethan-1,2-diol	should be **ethane-1,2-diol**
2-methpropan-2-ol	should be **2-methylpropan-2-ol**
2-methylbutan-3-ol	should be **3-methylbutan-2-ol**
3-methylpentan	should be **3-methylpentane**
3-mythylpentane	should be **3-methylpentane**
3-methypentane	should be **3-methylpentane**
propanitrile	should be **propanenitrile**
aminethane	should be **ethylamine** (although aminoethane can gain credit)
2-methyl-3-bromobutane	should be **2-bromo-3-methylbutane**
3-bromo-2-methylbutane	should be **2-bromo-3-methylbutane**
3-methyl-2-bromobutane	should be **2-bromo-3-methylbutane**
2-methylbut-3-ene	should be **3-methylbut-1-ene**
difluorodichloromethane	should be **dichlorodifluoromethane**

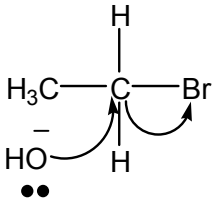
3.14 Organic reaction mechanisms

Curly arrows should originate either from a lone pair of electrons or from a bond.

The following representations should not gain credit **and will be penalised each time** within a clip.



For example, the following would score zero marks



When the curly arrow is showing the formation of a bond to an atom, the arrow can go directly to the relevant atom, alongside the relevant atom or **more than half-way** towards the relevant atom.

In free-radical substitution

- the absence of a radical dot should be penalised **once only** within a clip.
- the use of half-headed arrows is not required, but the use of double-headed arrows or the incorrect use of half-headed arrows in free-radical mechanisms should be penalised **once only** within a clip

The correct use of skeletal formulae in mechanisms is acceptable, but where a C-H bond breaks both the bond and the H must be drawn to gain credit.

Question	Marking guidance	Mark	Comments
01.1	Amount / number / proportion / percentage / fraction / moles (of molecules / particles/ atoms)	1	Penalise an incorrect qualification of the number eg NOT number of molecules with *E* greater than E_a
01.2	The most probable energy of the particles	1	
01.3	None / no effect / stays the same	1	
01.4	**M1** The peak of the new curve is displaced to the left and higher than the original **M2** All of the following needed • The new curve starts at the origin and should begin to separate from the original almost immediately • and the new curve only crosses the original curve once • and the total area under the new curve is approximately the same as the original • and an attempt has been made to draw the new curve correctly towards the axis below the original curve but not to touch the original curve	1 1	
01.5	Speeds up a reaction and is not changed (in chemical composition or amount)	1	Alters the rate of reaction Allow higher level answers eg lowers activation energy and provides alternative route

01.6	(decreasing the overall pressure) the rate of reaction decreases	1	
	because the collisions between reactant particles are less frequent (increasing the temperature)	1	
	the rate of reaction increases	1	
	more collisions between particles with E_a	1	Allow more particles have $E > E_a$

Question	Marking guidance	Mark	Comments
02.1	$N_2O_4 + 2N_2H_4 \rightarrow 4H_2O + 3N_2$	1	Ignore ss even if wrong Allow multiples
02.2	(N in N_2O_4) (+)4	1	
	(N in N_2H_4) –2	1	Sign needed
	The oxidation state of N (in N_2O_4) decreases and (in N_2H_4) increases (to zero)	1	Allow electrons are lost <u>and</u> gained
02.3	Concentration(s) (of reactants and products) remain(s) constant / stay(s) the same / remain(s) the same / do(es) not change	1	Accept [ ] for concentration **NOT** "equal concentrations" and **NOT** "concentration(s) is/are the same" **NOT** "amount"
	Forward rate = Reverse / backward rate	1	Ignore "dynamic" and ignore "speed" Ignore "closed system"
			It is possible to score both marks under the heading of a single feature

02.4	The (forward) reaction / to the right is endothermic or takes in / absorbs heat OR The reverse reaction / to the left is exothermic or gives out / releases heat **Next mark depends on correct first mark and must refer to temperature/heat** The equilibrium shifts / moves left to right (becomes darker) to oppose the increase in temperature	1 1	Accept the equilibrium shifts/moves to absorb the heat **OR** to lower the temperature **OR** to cool the reaction

02.5	There are fewer moles (of gas) on the left **OR** more moles (of gas) on the right. OR there is one mole (of gas) on the left and 2 moles on the right. **Second mark depends on correct first mark and must refer to pressure** The equilibrium shifts / moves right to left (becomes paler) to oppose the increase in pressure	1 1	Accept the equilibrium shifts/moves to lower the pressure

Question	Marking guidance	Mark	Comments
03.1	Hydrogen / H_2 $2H_2 + O_2 \rightarrow 2H_2O$	1 1	
03.2	Both white (solids)	1	Not same appearance
03.3	$Mg(OH)_2 + 2HCl \rightarrow MgCl_2 + 2H_2O$	1	Allow multiples
03.4	Add NaOH(aq) White ppt indicates Mg^{2+} Or No ppt indicates Sr^{2+} OR Add soluble sulfate eg Na_2SO_4(aq) Colourless soln indicates Mg^{2+} Or white ppt indicates Sr^{2+}	1 1	

Question	Marking guidance	Mark	Comments
04.1	(A) NaI / sodium iodide (B) NaBr / sodium bromide (C) NaCl / sodium chloride	1 1 1	 Allow NaF
04.2	$NaX + H_2SO_4 \rightarrow NaHSO_4 + HX$	1	Allow equations from NaCl or NaF Allow equations to form Na_2SO_4
04.3	NaI / sodium iodide	1	
04.4	$2H_2SO_4 + 2NaBr \rightarrow Na_2SO_4 + SO_2 + Br_2 + 2H_2O$ Or $3H_2SO_4 + 2NaBr \rightarrow 2NaHSO_4 + SO_2 + Br_2 + 2H_2O$	1	Allow $H_2SO_4 + 2Br^- + 2H^+ \rightarrow SO_2 + Br_2 + 2H_2O$ or $SO_4^{2-} + 2Br^- + 4H^+ \rightarrow SO_2 + Br_2 + 2H_2O$
04.5	I^- / iodide	1	

Question	Marking guidance	Mark	Comments
05.1	Average/mean mass of 1 atom (of an element) 1/12 mass of one atom of 12C ***OR*** Average/mean mass of atoms of an element 1/12 mass of one atom of 12C ***OR*** Average/mean mass of atoms of an element × 12 mass of one atom of 12C ***OR*** (Average) mass of one mole of atoms 1/12 mass of one mole of 12C ***OR*** (Weighted) average mass of all the isotopes 1/12 mass of one atom of 12C ***OR*** Average mass of an atom/all isotopes (compared to C-12) on a scale in which an atom of C-12 has a mass of 12	2	If moles and atoms mixed, max = 1 Mark top and bottom line independently. All key terms must be present for each mark.
05.2	$\frac{(58 \times 19) + (60 \times 8) + (62 \times 3)}{30}$ 58.9	1 1	
05.3	To stop the ions produced colliding with air or other particles/ so that there are only the ions produced from the sample present	1	**Not** just to remove air
05.4	$(1s^2)2s^2 2p^6 3s^2 3p^6 3d^8 4s^1$	1	

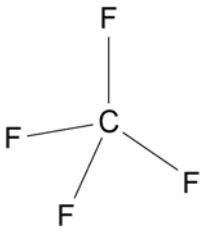
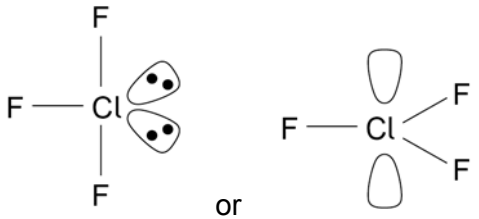
05.5	9.63×10^{-26} (kg)	1	

05.6	$KE = \frac{1}{2} m(\frac{d}{t})^2$	1	If 7.37×10^{-25} used
	$= \frac{9.63 \times 10^{-26}}{2} \times \frac{1.80^2}{(5.82 \times 10^{-7})^2}$	1	$= \frac{7.37 \times 10^{-25}}{2} \times \frac{1.80^2}{(5.82 \times 10^{-7})^2}$
	4.61×10^{-13} (J)	1	3.52×10^{-12} (J)

Question	Marking guidance	Mark	Comments
06.1	Amount of calcium nitrate = 4.38/164.1 = 0.0267 mol	1	Allow 0.027 but penalise 0.026 and 0.03 for M1
	Amount of gas products = $0.0267 \times \frac{5}{2}$ = 0.0667 mol	1	If M2 incorrect Max 3 (M1 M3 and M4)
	T = 823 K and P = 100 000 Pa	1	M3 allow 550 + 273 for temperature
	V = nRT/P	1	M4 allow PV = nRT or any correct rearrangement.
	$\left(\frac{0.0667 \times 8.31 \times 823}{100\,000}\right) = 4.56 \times 10^{-3}\ (m^3)$	1	M5 allow 4.56×10^{-3} - $4.58 \times 10^{-3}\ (m^3)$
06.2	Difficult to separate two gases	1	
06.3	(1) $Ca_3(PO_4)_2(s) + 6\ HNO_3(aq) \rightarrow 2\ H_3PO_4(aq) + 3\ Ca(NO_3)_2(aq)$	1	

06.4	Amount of calcium carbonate = 2.1×10^{-3} mol and amount of nitric acid = 0.0125 mol 1:2 ratio so limiting reagent is 2.1×10^{-3} mol calcium carbonate Amount of calcium nitrate formed = 2.1×10^{-3} mol Concentration of calcium nitrate = $2.1 \times 10^{-3} / 25 \times 10^{-3}$ = <u>0.084</u> mol dm^{-3}	1 1 1	 Allow answer to 2sf or more

Question	Marking guidance	Mark	Comments
07.1	F/ fluorine	1	
07.2	Ne	1	If not neon then CE = 0 / 3. But if Ar chosen, lose M1 and allow M2+M3
			Allow similar shielding
	Large(st) number of protons / large(st) nuclear charge	1	Ignore smallest atomic radius
	Same amount of shielding / same number of shells / same number of energy levels	1	Allow similar shielding
07.3	$O(g) \rightarrow O^{+}(g) + e^{(-)}$	1	
07.4	C	1	If not carbon then CE = 0 / 3
	covalent (bonds)	1	M3 dependent on correct M2
	Strong or many of the (covalent) bonds need to be broken / needs a lot of energy to break the (covalent) bonds	1	Mention of ions would lose M2 and M3 Ignore IMF

07.5	  Shape = tetrahedral	1 1 1	

07.6	van der Waals/vdw forces between the molecules that are very weak (and easily broken by heat)	1 1	not breaking covalent bonds

07.7	C Cl F 11.8% (M1) 69.6 % 18.6% 12 35.5 19 (=0.983 =1.961 =0.980) 1 2 1 OR CCl_2F CCl_2F has M_r = 102 MF = $C_2Cl_4F_2$	1 1 1 1	M1 11.8% M2 dividing by correct A_r M3 correct simplest ratio or CCl_2F M4 correct MF Allow alternative methods