

Please check the examination details below before entering your candidate information

Candidate surname

Other names

Centre Number

Candidate Number

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Pearson Edexcel International Advanced Level

Friday 20 October 2023

Morning (Time: 1 hour 45 minutes)

Paper
reference

WCH14/01

Chemistry

International Advanced Level

**UNIT 4: Rates, Equilibria and Further Organic
Chemistry**

You must have:

Scientific calculator, Data Booklet, ruler

Total Marks

Instructions

- Use **black** ink or ball-point pen.
- If pencil is used for diagrams/sketches/graphs it must be dark (HB or B).
- **Fill in the boxes** at the top of this page with your name, centre number and candidate number.
- Answer **all** questions.
- Answer the questions in the spaces provided
– *there may be more space than you need.*

Information

- The total mark for this paper is 90.
- The marks for **each** question are shown in brackets
– *use this as a guide as to how much time to spend on each question.*
- In the question marked with an **asterisk (*)**, marks will be awarded for your ability to structure your answer logically, showing how the points that you make are related or follow on from each other where appropriate.
- A Periodic Table is printed on the back cover of this paper.

Advice

- Read each question carefully before you start to answer it.
- Show all your working in calculations and include units where appropriate.
- Try to answer every question.
- Check your answers if you have time at the end.

Turn over ►

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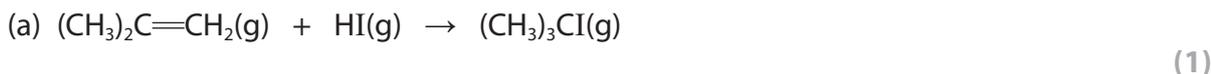
SECTION A

Answer ALL the questions in this section.

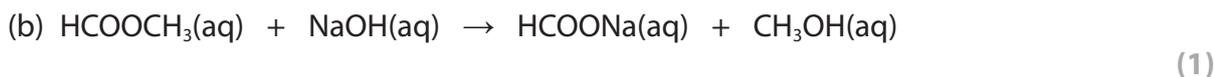
You should aim to spend no more than 20 minutes on this section.

For each question, select one answer from A to D and put a cross in the box ☒.
If you change your mind, put a line through the box ☒ and then mark your new answer with a cross ☒.

1 Which method would be most suitable to investigate the kinetics of the reactions shown?



- A colorimetry
 B measurement of change in volume
 C measurement of change in mass
 D quenching with ice-cold water followed by titrating with acid



- A colorimetry
 B measurement of change in volume
 C measurement of change in mass
 D quenching with ice-cold water followed by titrating with acid

(Total for Question 1 = 2 marks)

2 The equation for the reaction of bromate(V) ions with bromide ions in acid solution is shown.



The rate equation for the reaction is

$$\text{rate} = k[\text{BrO}_3^-][\text{Br}^-][\text{H}^+]^2$$

The concentrations of all the reactants are doubled.
By what factor does the rate of reaction increase?

- A 2
 B 4
 C 8
 D 16

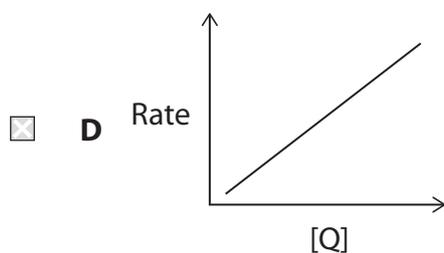
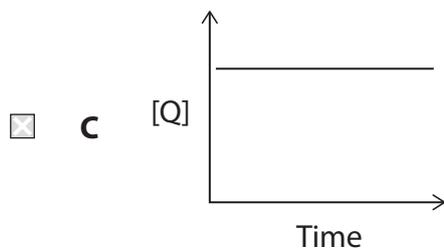
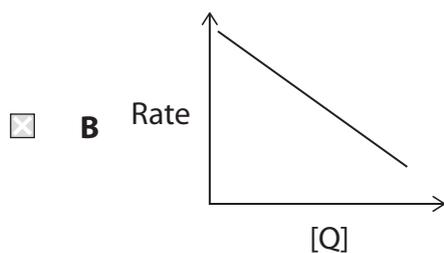
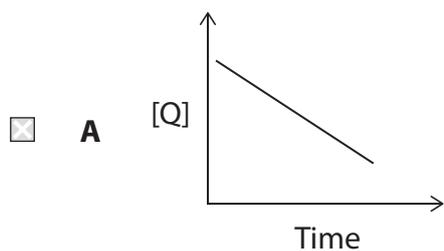
(Total for Question 2 = 1 mark)



3 This question is about rates of chemical reactions.

(a) Which graph shows a reaction that is zero order with respect to reactant Q?

(1)



(b) The equation for a gas phase reaction is shown.



The reaction is first order.

When the initial pressure of A is 2 atm the half-life of the reaction is 20 s.

What is the half-life of the reaction when the initial pressure of A is 4 atm?

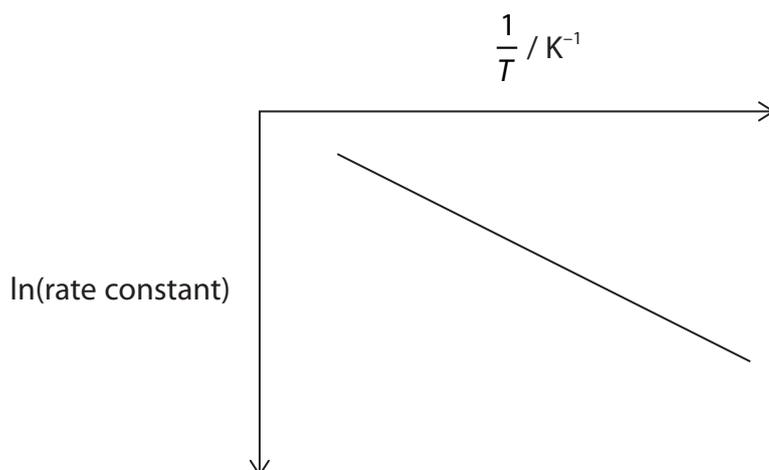
(1)

- A** 10 s
- B** 20 s
- C** 40 s
- D** 400 s

(Total for Question 3 = 2 marks)



- 4 The graph shown can be used to obtain a value for the activation energy, E_a , of a reaction.



The activation energy is related to the rate constant by the equation

$$\ln(\text{rate constant}) = -\frac{E_a}{R} \times \frac{1}{T} + \text{constant}$$

From the graph, the E_a is equal to

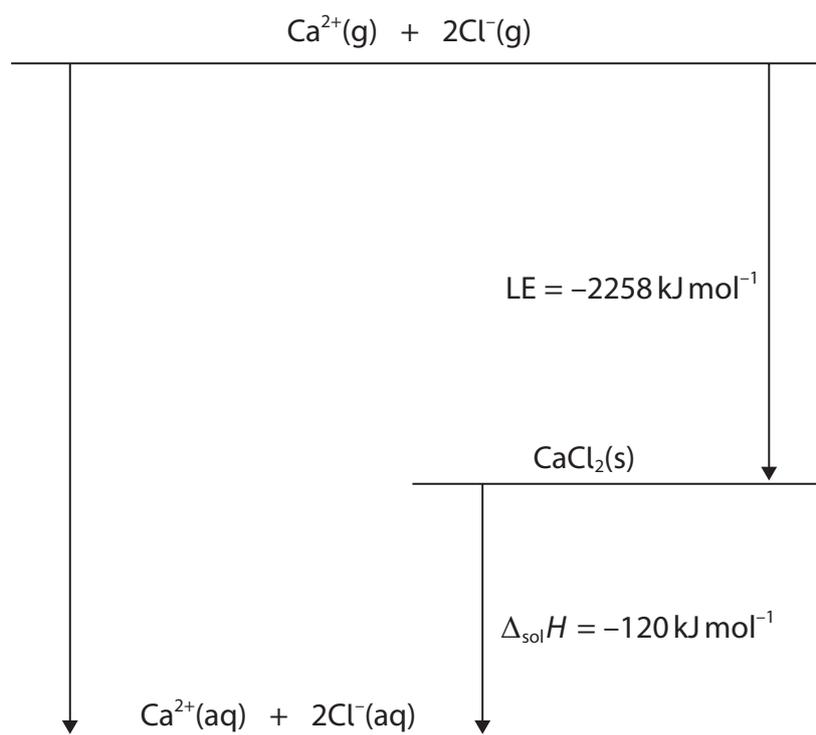
- A $\frac{(-\text{gradient})}{R}$
- B $\frac{(-\text{gradient})}{RT}$
- C $(-\text{gradient}) \times R$
- D $(-\text{gradient}) \times RT$

(Total for Question 4 = 1 mark)

Use this space for any rough working. Anything you write in this space will gain no credit.



5 The energy cycle for dissolving calcium chloride in water is shown.



(a) The enthalpy change of hydration of the calcium ion is $-1650 \text{ kJ mol}^{-1}$.

What is the enthalpy change of hydration, in kJ mol^{-1} , of the chloride ion?

(1)

- A -728
- B -364
- C $+364$
- D $+728$

(b) Why is the enthalpy change of hydration for magnesium ions more exothermic than that for calcium ions?

(1)

- A magnesium ions have a larger radius
- B magnesium ions form stronger ionic bonds
- C magnesium ions have a higher charge density
- D magnesium has higher first and second ionisation energies

(Total for Question 5 = 2 marks)



6 Heating copper(II) nitrate results in the equilibrium shown.

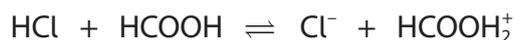


Which is the expression for K_p ?

- A $K_p = \frac{2(p\text{CuO}) \times 4(p\text{NO}_2) \times (p\text{O}_2)}{2(p\text{Cu}(\text{NO}_3)_2)}$
- B $K_p = \frac{(p\text{CuO})^2 \times (p\text{NO}_2)^4 \times (p\text{O}_2)}{(p\text{Cu}(\text{NO}_3)_2)^2}$
- C $K_p = 4(p\text{NO}_2) \times (p\text{O}_2)$
- D $K_p = (p\text{NO}_2)^4 \times (p\text{O}_2)$

(Total for Question 6 = 1 mark)

7 When concentrated hydrochloric acid is added to methanoic acid, an acid-base reaction occurs.



What are the Brønsted-Lowry acid-base conjugate pairs in this equilibrium?

	Acid 1	Conjugate base of Acid 1	Acid 2	Conjugate base of Acid 2
<input type="checkbox"/> A	HCl	HCOOH	HCOOH ₂ ⁺	Cl ⁻
<input type="checkbox"/> B	HCl	HCOOH ₂ ⁺	HCOOH	HCOOH ₂ ⁺
<input type="checkbox"/> C	HCl	Cl ⁻	HCOOH	HCOOH ₂ ⁺
<input type="checkbox"/> D	HCl	Cl ⁻	HCOOH ₂ ⁺	HCOOH

(Total for Question 7 = 1 mark)

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8 At 100 °C, pure water has a pH of 6, but at 25 °C it has a pH of 7.

This is because

- A the dissociation of water is endothermic, so the concentration of hydrogen ions is lower at 100 °C than at 25 °C
- B the dissociation of water is exothermic, so the concentration of hydrogen ions is lower at 100 °C than at 25 °C
- C the dissociation of water is endothermic, so the concentration of hydrogen ions is higher at 100 °C than at 25 °C
- D the dissociation of water is exothermic, so the concentration of hydrogen ions is higher at 100 °C than at 25 °C

(Total for Question 8 = 1 mark)

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9 Four beakers each contain equal volumes of four solutions.

Beaker 1 contains $0.01 \text{ mol dm}^{-3} \text{ NaOH}$

Beaker 2 contains $0.01 \text{ mol dm}^{-3} \text{ NH}_3$

Beaker 3 contains $0.1 \text{ mol dm}^{-3} \text{ NaOH}$

Beaker 4 contains $0.1 \text{ mol dm}^{-3} \text{ Ba(OH)}_2$

The pH of the four solutions was measured.

Which of the following gives the order of **decreasing** pH?

- A 1, 2, 3, 4
- B 2, 1, 3, 4
- C 3, 4, 1, 2
- D 4, 3, 1, 2

(Total for Question 9 = 1 mark)

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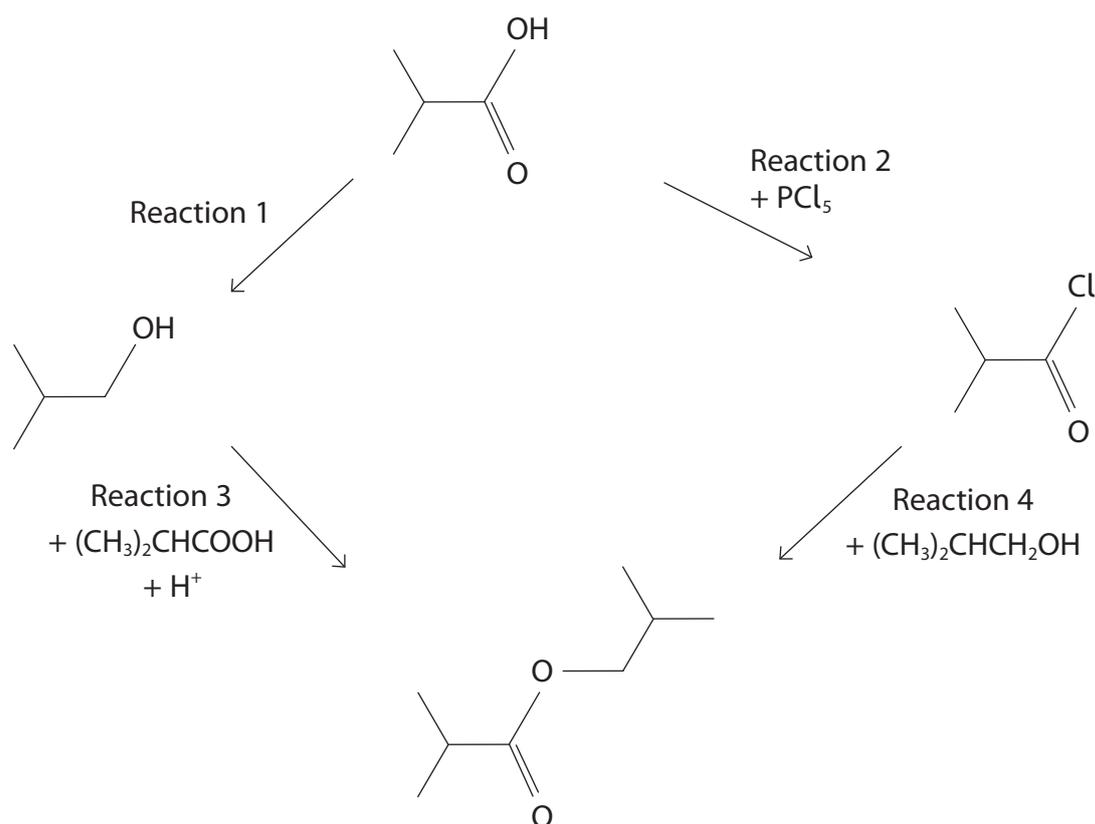
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10 This question is about the reaction scheme shown.



(a) Which reagent is required for Reaction 1?

(1)

- A acidified potassium dichromate(VI)
- B concentrated sulfuric acid
- C hydrogen with a nickel catalyst
- D lithium tetrahydridoaluminate(III)

(b) Reaction 1 has a yield of 90%.

What mass of 2-methylpropanoic acid is required to produce 6.66 g of 2-methylpropan-1-ol?

[*M_r* values 2-methylpropanoic acid = 88 2-methylpropan-1-ol = 74]

(1)

- A 6.22 g
- B 7.13 g
- C 7.92 g
- D 8.80 g



(c) Which condition is essential for Reaction 2?

(1)

- A acid catalyst
- B anhydrous
- C high temperature
- D ether solvent

(d) Reactions 3 and 4 produce the same ester.

Which is an advantage of using Reaction 4 to produce this ester in the laboratory?

(1)

- A it can be carried out at room temperature
- B it requires a catalyst
- C the atom economy is higher
- D hydrogen chloride is a product

(Total for Question 10 = 4 marks)

11 Which diagram shows two repeat units of the polymer formed by the polymerisation of propane-1,3-diol and benzene-1,4-dicarboxylic acid?

- A
- B
- C
- D

(Total for Question 11 = 1 mark)

12 Polymer waste may be disposed of by incineration.

At a plant carrying out this process with poly(propene), a waste gas was produced that was thought to be either propane or carbon dioxide. These compounds have the same molecular ion peak in a low resolution mass spectrometer but can be separated at high resolution.

Atom	Relative atomic mass, A_r
^1H	1.0079
^{12}C	12.0000
^{16}O	15.9949

Which are the correct relative molecular masses?

	propane	carbon dioxide
<input type="checkbox"/> A	29.0395	27.9949
<input type="checkbox"/> B	27.9949	29.0395
<input type="checkbox"/> C	43.9898	44.0632
<input type="checkbox"/> D	44.0632	43.9898

(Total for Question 12 = 1 mark)

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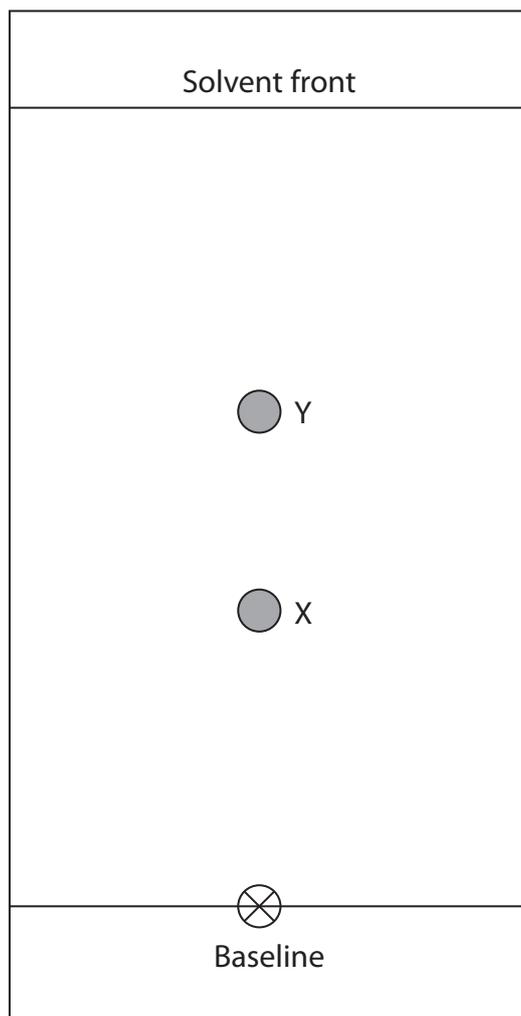
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P 7 5 0 6 8 A 0 1 3 3 6

13 A thin-layer chromatogram of a mixture of two compounds, X and Y, is shown.



(a) What is the R_f value for compound X?

(1)

- A 0.33
- B 0.38
- C 0.60
- D 0.62



(b) Y travelled further than X. How do the attractions between Y and X to the stationary and mobile phases compare?

(1)

	Attraction between Y and the stationary phase	Attraction between Y and the mobile phase
<input type="checkbox"/> A	is stronger than the attraction between X and the stationary phase	is stronger than the attraction between X and the mobile phase
<input type="checkbox"/> B	is stronger than the attraction between X and the stationary phase	is weaker than the attraction between X and the mobile phase
<input type="checkbox"/> C	is weaker than the attraction between X and the stationary phase	is stronger than the attraction between X and the mobile phase
<input type="checkbox"/> D	is weaker than the attraction between X and the stationary phase	is weaker than the attraction between X and the mobile phase

(Total for Question 13 = 2 marks)

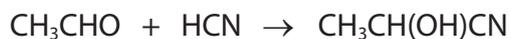
TOTAL FOR SECTION A = 20 MARKS



SECTION B

Answer ALL the questions. Write your answers in the spaces provided.

- 14** This question is about the nucleophilic addition reaction between ethanal and hydrogen cyanide in the presence of potassium cyanide as a catalyst. The equation for the reaction is shown.

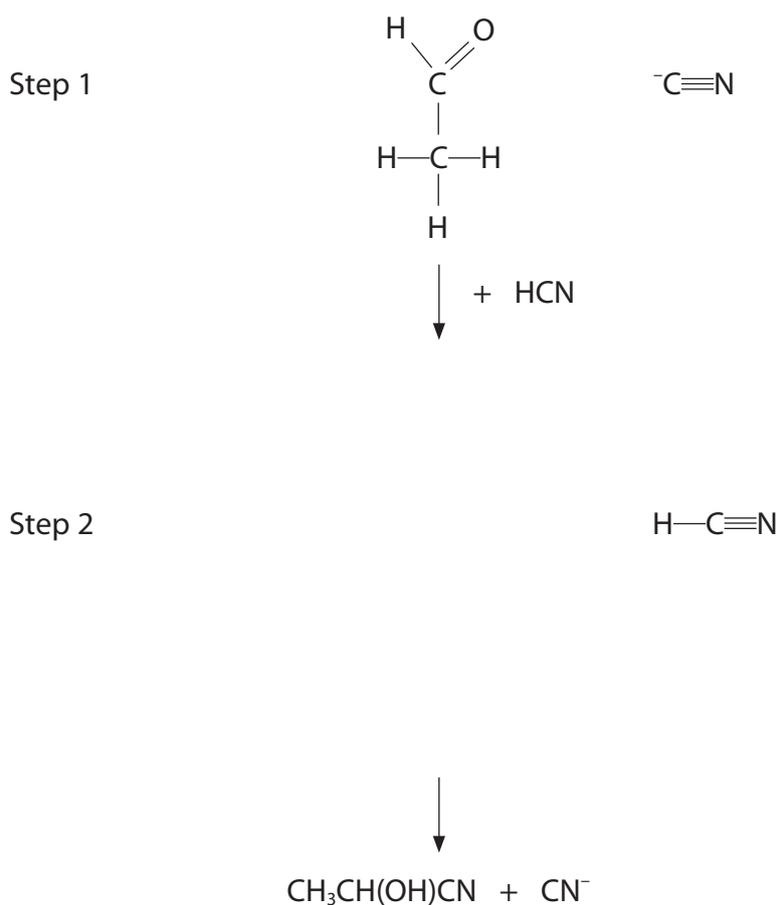


- (a) (i) Give the IUPAC name of the organic product.

(1)

- (ii) Complete the mechanism for this two-step reaction, showing the structure of the intermediate and including curly arrows, and relevant lone pairs and dipoles.

(4)



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(a) (iii) Justify the use of the term 'nucleophilic addition' to describe the mechanism of this reaction.

(2)

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(b) Explain why the product of the reaction is **not** optically active even though it contains a chiral carbon atom.

(2)

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(Total for Question 14 = 9 marks)



- 15 A reaction vessel contained nitrogen monoxide and oxygen in a 2:1 molar ratio. The mixture was allowed to come to equilibrium forming nitrogen dioxide. The equation for the reaction is shown.



The volume of the vessel was 15 dm^3 and the reaction was carried out at a constant temperature and at a pressure of $200\,000 \text{ Pa}$.

At equilibrium there was a total of 0.69625 mol of gas in the reaction vessel and the mass of oxygen was 7.000 g .

- (a) (i) Calculate the number of moles of each substance at equilibrium.

(3)

- (ii) Calculate the value of K_c under these conditions. Include units in your answer.

(4)



(b) Calculate the temperature, in K , of the reaction mixture at equilibrium under these conditions.

Use the equation $pV = nRT$ and the data at the start of the question.

(3)

(c) Under a different set of conditions, the reaction was carried out to find the initial rate of reaction.

Experiment number	Initial $[\text{NO}] / \text{mol dm}^{-3}$	Initial $[\text{O}_2] / \text{mol dm}^{-3}$	Initial rate / $\text{mol dm}^{-3} \text{s}^{-1}$
1	6.50×10^{-2}	1.25×10^{-2}	6.87×10^{-4}

The rate equation for this reaction is

$$\text{rate} = k[\text{NO}]^2[\text{O}_2]$$

(i) Give the reason why colorimetry can be used to monitor the rate of the reaction.

(1)



- (ii) Use the data from Experiment 1 to find the value of the rate constant, k .
Include units in your answer.

(2)

- (iii) State why the reaction is unlikely to proceed in a single step.

(1)

- (iv) A student proposed the mechanism shown for this reaction.



Justify whether or not this mechanism is consistent with the overall equation for the reaction **and** with the rate equation.

(2)

(Total for Question 15 = 16 marks)



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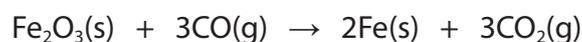
16 Iron, lead and zinc can be extracted using a blast furnace but aluminium cannot.

In a blast furnace, iron is extracted from iron(III) oxide, Fe_2O_3 , at high temperature.

Some standard enthalpy changes of formation, $\Delta_f H^\ominus$, and standard molar entropies, S^\ominus , are shown.

Substance	Al(s)	$\text{Al}_2\text{O}_3(\text{s})$	$\text{CO}(\text{g})$	$\text{CO}_2(\text{g})$	Fe(s)	$\text{Fe}_2\text{O}_3(\text{s})$
$\Delta_f H^\ominus / \text{kJ mol}^{-1}$	0	-1676	-111	-394	0	-824
$S^\ominus / \text{JK}^{-1} \text{mol}^{-1}$	28.3	50.9	197.6	213.6	27.3	87.4

(a) The main reaction occurring in the blast furnace to form iron is shown.



This reaction is feasible at all temperatures.

(i) Calculate the standard entropy change of the system for this reaction.

(3)

(ii) Calculate the standard enthalpy change for this reaction.

(3)



(iii) Explain how your answers to (a)(i) and (a)(ii) show that this reaction is feasible at all temperatures.

(3)

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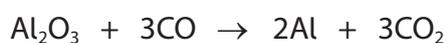
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(b) The main reduction reaction of aluminium oxide in a blast furnace is shown.



(i) Calculate the temperature at which this reaction becomes feasible.

(4)

(ii) Suggest why aluminium is not extracted from its oxide using a blast furnace.

(1)

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(Total for Question 16 = 14 marks)

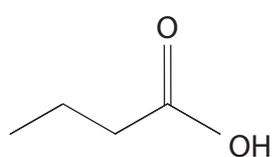
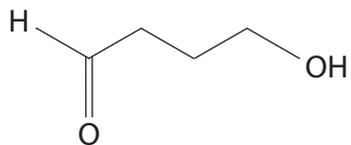
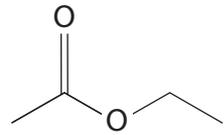
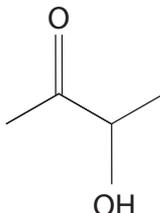
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17 This question is about four isomers with the molecular formula $C_4H_8O_2$.

Name	Skeletal formula
butanoic acid	
4-hydroxybutanal	
ethyl ethanoate	
3-hydroxybutanone	

- *(a) These four isomers can be identified using three **chemical** tests. Each test gives a positive result for only one isomer. All three tests give a negative result for the fourth isomer.

Deduce the three **chemical** tests required. For each test

- identify the reagent
- give the positive observation in each test
- identify the functional group of the isomer that gives the positive result.

(6)

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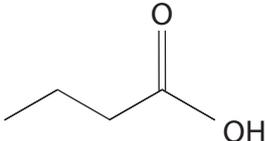
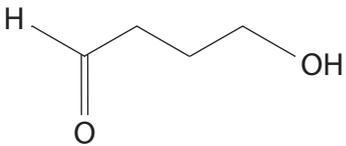
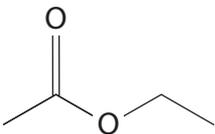
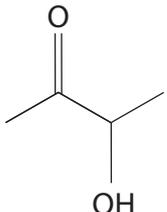
(b) The four isomers can also be distinguished using NMR spectroscopy.

- (i) State why it is **not** possible to distinguish any of these isomers from the number of peaks in their ^{13}C NMR spectra.

(1)

- (ii) Complete the table below to give the number of peaks in a **low** resolution ^1H NMR spectrum of each isomer.

(3)

Name	Skeletal formula	Number of peaks
butanoic acid		
4-hydroxybutanal		
ethyl ethanoate		
3-hydroxybutanone		



(iii) One of the ^1H NMR spectra has a peak with a chemical shift, δ , greater than 10 ppm.
Identify the isomer, and the proton environment responsible for this peak.

(1)

(iv) The high resolution ^1H NMR spectrum of one of the isomers contains a multiplet of five peaks (a quintet).
Explain, with reference to the structure of one of the isomers, the presence of the quintet.

(2)

(Total for Question 17 = 13 marks)

TOTAL FOR SECTION B = 52 MARKS

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SECTION C

Answer ALL the questions. Write your answers in the spaces provided.

18 Pentanoic acid, C_4H_9COOH , is a carboxylic acid with $K_a = 1.38 \times 10^{-5} \text{ mol dm}^{-3}$.

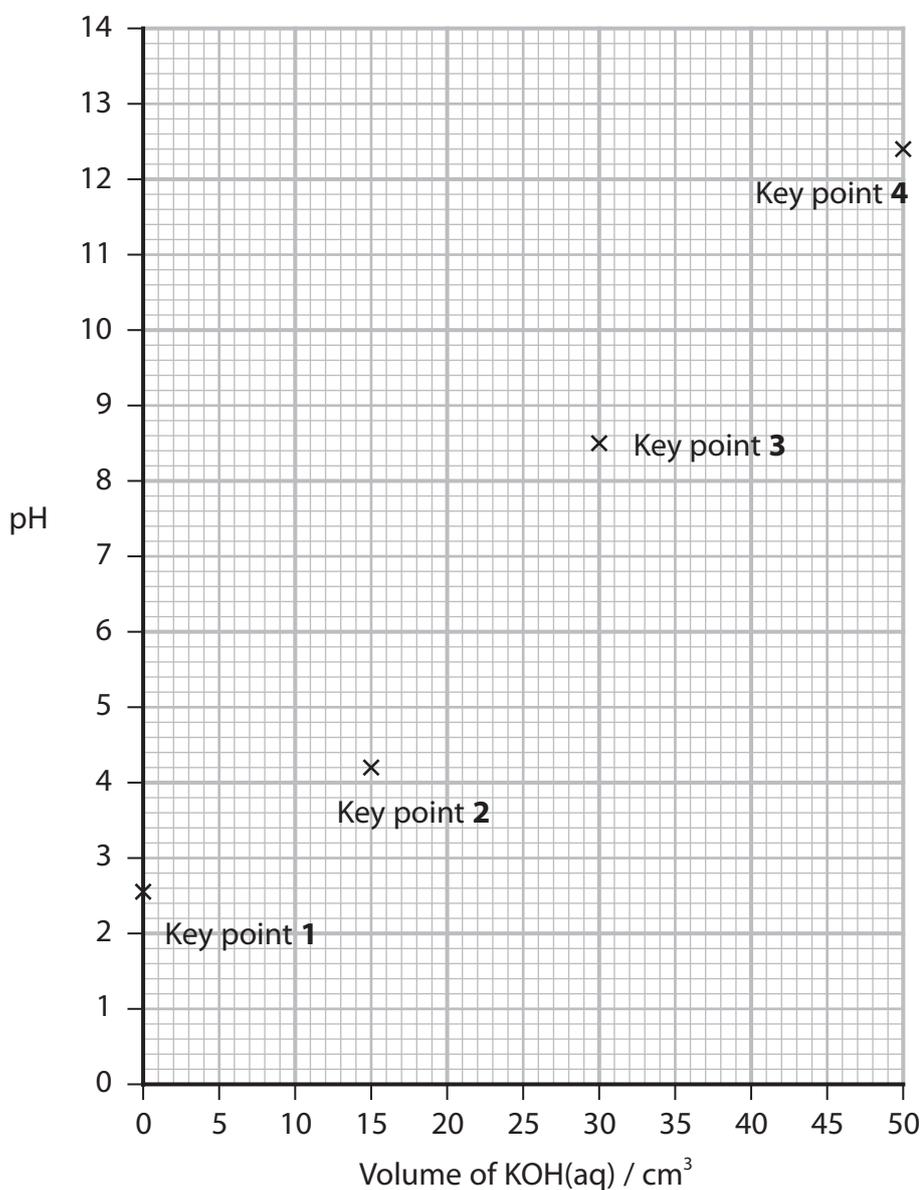
A student was asked to titrate 25.0 cm^3 of $0.120 \text{ mol dm}^{-3}$ pentanoic acid solution with $0.100 \text{ mol dm}^{-3}$ potassium hydroxide solution.

The equation for the reaction is shown.



The teacher asked the student to sketch a graph showing the expected changes in pH as potassium hydroxide solution is added.

The student first identified four key points for the sketch graph, using four calculations and one chemical equation. These points are shown.



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(a) (i) At Key point **1**, before any potassium hydroxide has been added, the pH = 2.9.
Use a calculation to justify this value.

(3)

(ii) The student deduced that the pH at Key point **4** must be less than 13.
Use a calculation to justify the student's deduction.
You are **not** required to calculate the exact pH of the solution.

(2)

(iii) The student deduced that the neutralisation point of the graph at Key point **3**
would be at 30 cm³.
Use a calculation to justify the student's deduction.

(2)



(iv) The student deduced that when equal amounts of pentanoic acid and potassium hydroxide have been mixed at Key point **3**, the pH would be greater than 7.

Justify the student's deduction.

(1)

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(v) The student deduced that at Key point **2**, 15.0 cm³ of potassium hydroxide would have been added and the pH would be 4.9.

Justify the student's deductions using a calculation.

(2)

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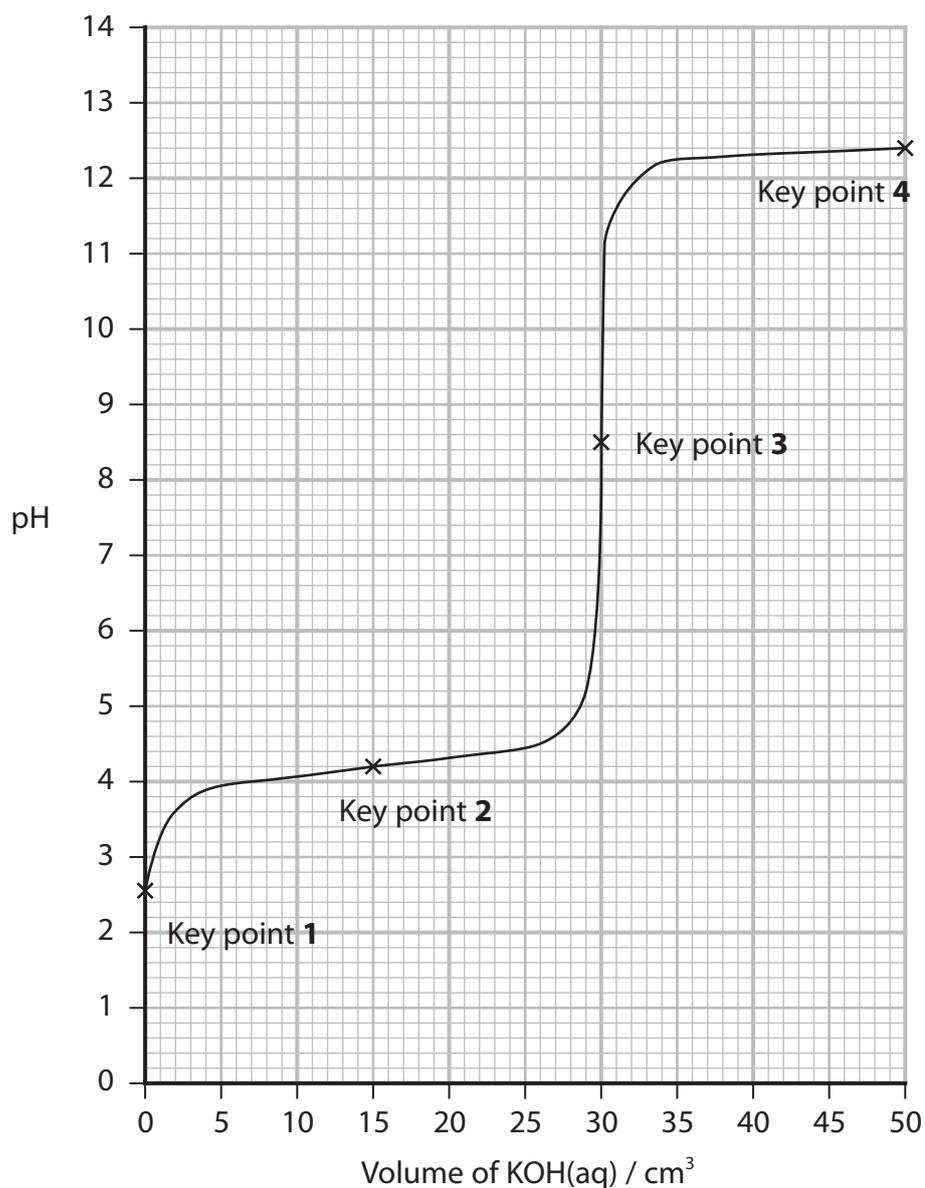
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(b) The student then completed the sketch graph by linking these four points to show how the pH would change as potassium hydroxide was added.



Explain why the student drew the line so that in the section of the graph between 5 cm³ and 25 cm³ the pH changes very little as the potassium hydroxide is added. No calculation is required.

(3)

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- (c) (i) The student carried out the titration using methyl orange as an indicator. The experiment did not give a satisfactory colour change at the neutralisation point, and the student continued adding potassium hydroxide until 50 cm^3 had been added.

Describe the colour of the solution during this titration experiment with particular reference to the key points on the sketch graph.

Use the Data Booklet.

(3)

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- (ii) The student found a bottle labelled 'Acid-base indicator solution' in the store cupboard. The student, having checked with the teacher, used this as the indicator for the titration.

The initial yellow solution turned green at the neutralisation point of the titration.

Suggest the identity of the indicator solution including a justification for the colour observed at the neutralisation point.

Use the Data Booklet.

(2)

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(Total for Question 18 = 18 marks)

TOTAL FOR SECTION C = 18 MARKS
TOTAL FOR PAPER = 90 MARKS



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P 7 5 0 6 8 A 0 3 5 3 6

The Periodic Table of Elements

1 2 3 4 5 6 7 0 (8) (18)

1.0	H	hydrogen	1
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Key

relative atomic mass
atomic symbol
name
atomic (proton) number

(1) (2)

6.9	Li	lithium	3
9.0	Be	beryllium	4
23.0	Na	sodium	11
24.3	Mg	magnesium	12

39.1	K	potassium	19
85.5	Rb	rubidium	37
132.9	Cs	caesium	55

45.0	Sc	scandium	21
47.9	Ti	titanium	22
50.9	V	vanadium	23
52.0	Cr	chromium	24
54.9	Mn	manganese	25
55.8	Fe	iron	26
58.9	Co	cobalt	27
58.7	Ni	nickel	28
63.5	Cu	copper	29
65.4	Zn	zinc	30
69.7	Ga	gallium	31
72.6	Ge	germanium	32
74.9	As	arsenic	33
79.0	Se	selenium	34
83.8	Kr	krypton	36

88.9	Y	yttrium	39
87.6	Sr	strontium	38
138.9	Ba	barium	56
178.5	Hf	hafnium	72
180.9	Ta	tantalum	73
186.2	Re	rhenium	75
190.2	Os	osmium	76
192.2	Ir	iridium	77
195.1	Pt	platinum	78
197.0	Au	gold	79
200.6	Hg	mercury	80
204.4	Tl	thallium	81
207.2	Pb	lead	82
209.0	Bi	bismuth	83
212.0	Po	polonium	84
210.0	At	astatine	85
222.0	Rn	radon	86

10.8	B	boron	5
12.0	C	carbon	6
14.0	N	nitrogen	7
16.0	O	oxygen	8
19.0	F	fluorine	9
20.2	Ne	neon	10
27.0	Al	aluminium	13
28.1	Si	silicon	14
31.0	P	phosphorus	15
32.1	S	sulfur	16
35.5	Cl	chlorine	17
39.9	Ar	argon	18

112.4	Cd	cadmium	48
114.8	In	indium	49
118.7	Sn	tin	50
121.8	Sb	antimony	51
127.6	Te	tellurium	52
127.6	Te	tellurium	52

112.4	Cd	cadmium	48
114.8	In	indium	49
118.7	Sn	tin	50
121.8	Sb	antimony	51
127.6	Te	tellurium	52
127.6	Te	tellurium	52

4.0	He	helium	2
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Elements with atomic numbers 112-116 have been reported but not fully authenticated

* Lanthanide series
* Actinide series

140	Ce	cerium	58
141	Pr	praseodymium	59
144	Nd	neodymium	60
147	Pm	promethium	61
150	Sm	samarium	62
152	Eu	europium	63
157	Gd	gadolinium	64
159	Tb	terbium	65
163	Dy	dysprosium	66
165	Ho	holmium	67
167	Er	erbium	68
169	Tm	thulium	69
173	Yb	ytterbium	70
175	Lu	lutetium	71
232	Th	thorium	90
231	Pa	protactinium	91
238	U	uranium	92
237	Np	neptunium	93
242	Pu	plutonium	94
243	Am	americium	95
247	Cm	curium	96
245	Bk	berkelium	97
251	Cf	californium	98
254	Es	einsteinium	99
253	Fm	fermium	100
256	Md	mendelevium	101
254	No	nobelium	102
257	Lr	lawrencium	103

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