Question Number	Answer	Mark
3	**The only correct answer is A** $((0.5 \times 436 + 0.5 \times 242) - 431)$ *B is not correct because the bond enthalpies of the reactants have been subtracted from the bond enthalpy of the product and this is for the formation of two moles of HCl* *C is not correct because the bond enthalpies of the reactants have been subtracted from the bond enthalpy of the product* *D is not correct because this is for the formation of two moles of HCl*	(1)

Question Number	Answer	Mark
2	**The only correct answer is A** ($\frac{1}{2}Br_2(l) \rightarrow Br(g)$) *B is not correct because bromine is a liquid in its standard state* *C is not correct because this shows the formation of two moles of gaseous bromine atoms* *D is not correct because bromine is a liquid in its standard state and this shows the formation of two moles of gaseous bromine atoms*	

Question Number	Answer
6(a)	**The only correct answer is D** $\left(\dfrac{-\Delta T \times 4.2 \times 50}{2.5 \times 10^{-3}} \right)$ *A is not correct because it has the wrong sign, and the wrong number of moles* *B is not correct because it has the wrong sign and the incorrect mass of solution* *C is not correct because it has the wrong number of moles*

Question Number	Answer
6(b)	**The only correct answer is D** ($25\ cm^3$ pipette) *A is not correct because the burette has to be read twice so the % uncertainty is 0.4%* *B is not correct because the % uncertainty is 2 %* *C is not correct because the % uncertainty is 0.4 %*